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MIXED REVIEW SHORT ANSWER

Answer the following questions in the space provided. 1. Given the following

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Equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$

4 a. What is the value of the coefficient x in this equation? 40.07 g/mol

b. What is the molar mass of C_3H_4 ? 2 mol O_2 :1 mol H_2O

c. What is the mole ratio of O_2 to H_2O ?

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of hydrogen peroxide (in aqueous solution)
decompose to produce two molecules of
liquid water and one molecule of oxygen

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SECTION 3 PROBLEMS Write the

answer on the line to the left. Show all your work in the space provided. 1. _____

The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H

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Stoichiometry b. Theoretically, how many moles of NH_3 will be produced?

PROBLEMS Write the answer on the line to the left, Show all your work in the space provided. 1 88% The actual yield of a reaction is 22 g and the theoretical yield is

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25 g. Calculate the percentage yield. 2. 6.0 mol of N_2 are mixed with 12.0 mol of H_2 according to the ...

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PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. The following equation represents a laboratory preparation for oxygen gas: ... CHAPTER 9 REVIEW ...

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Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$

This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce

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two molecules of liquid water and one molecule of oxygen gas.

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Introduction to Stoichiometry.

Composition Stoichiometry - deals with
mass relationships of elements in

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Compounds Reaction Stoichiometry -

Involves mass relationships between reactants and products in a chemical

reaction. I. Reaction Stoichiometry

Problems A. Four problem Types, One Common Solution.

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Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ a. What is the value of the coefficient x in this equation? 40.07

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g/mol b. What is the molar

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Answer the following questions in the space provided. 1. Given the following

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Equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ 4 a.

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