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1 a) The relative atomic mass of an element is the weighted average mass of the isotopes of the element. (1)

It is measured on a scale on which a carbon-12 atom has a mass of

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units. (1). b) The
total mass of

$$100 \text{ atoms} = (75 \times 35) + (25 \times 37) = 3550 \text{ (1)}$$

The average mass
of 1 atom =
 $3550/100 = 35.5$
(1)

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a) melting b)

freezing c)

subliming /

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sublimation 2.

a) Note: Solids

should have

regularly packed

particles

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